

PCT

WORLD INTELLECTUAL PROPERTY ORGANIZATION
International Bureau



INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification ⁶ : B01J 31/18, 31/12, 31/16, C07F 7/10, 7/28		A1	(11) International Publication Number: WO 96/27439
			(43) International Publication Date: 12 September 1996 (12.09.96)
(21) International Application Number: PCT/EP96/01011		(81) Designated States: AL, AM, AT, AU, AZ, BB, BG, BR, BY, CA, CH, CN, CZ, DE, DK, EE, ES, FI, GB, GE, HU, IS, JP, KE, KG, KP, KR, KZ, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, TJ, TM, TR, TT, UA, UG, UZ, VN, ARIPO patent (KE, LS, MW, SD, SZ, UG), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG).	
(22) International Filing Date: 6 March 1996 (06.03.96)			
(30) Priority Data: 95200570.0 8 March 1995 (08.03.95) EP (34) Countries for which the regional or international application was filed: GB et al.			
(71) Applicant (for all designated States except CA): SHELL INTERNATIONALE RESEARCH MAATSCHAPPIJ B.V. [NL/NL]; Carel van Bylandtlaan 30, NL-2596 HR The Hague (NL).		<p>Published</p> <p>With international search report.</p> <p>Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.</p>	
(71) Applicant (for CA only): SHELL CANADA LIMITED [CA/CA]; 400 4th Avenue S.W., Calgary, Alberta T2P 2H5 (CA).			
(72) Inventors: HORTON, Andrew, David; Badhuisweg 3, NL-1031 CM Amsterdam (NL). DE WITTH, Jan; Badhuisweg 3, NL-1031 CM Amsterdam (NL).			
(54) Title: BRIDGED BIS-AMIDO GROUP 4 METAL COMPOUNDS IN A CATALYST COMPOSITION			
(57) Abstract			
<p>Provided is a catalyst composition comprising a bis-amide compound represented by formula (II), wherein M is zirconium, hafnium or titanium; each N is a three coordinate nitrogen atom; each Y is Si, Ge or Sn; each X is, independently, a hydride or R; each R is, independently, a hydrocarbyl one or more carbon atoms of which may be substituted by an element selected from Si, O, P, N and S; Q is a neutral Lewis base and m is a number from 0 to 2; or a dimer thereof and a second component. Preferably, the second component is a compound which is capable of providing a bulky and labile anion [A]⁻, which anion is substantially non-coordinating under the reaction conditions and contains at least one boron atom. Also provided is a process for oligomerising or co-oligomerising alpha olefins in the presence of this catalyst composition.</p>			
<p style="text-align: right;">(II)</p>			

FOR THE PURPOSES OF INFORMATION ONLY

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AM	Armenia	GB	United Kingdom	MW	Malawi
AT	Austria	GE	Georgia	MX	Mexico
AU	Australia	GN	Guinea	NE	Niger
BB	Barbados	GR	Greece	NL	Netherlands
BE	Belgium	HU	Hungary	NO	Norway
BF	Burkina Faso	IE	Ireland	NZ	New Zealand
BG	Bulgaria	IT	Italy	PL	Poland
BJ	Benin	JP	Japan	PT	Portugal
BR	Brazil	KE	Kenya	RO	Romania
BY	Belarus	KG	Kyrgyzstan	RU	Russian Federation
CA	Canada	KP	Democratic People's Republic of Korea	SD	Sudan
CF	Central African Republic	KR	Republic of Korea	SE	Sweden
CG	Congo	KZ	Kazakhstan	SG	Singapore
CH	Switzerland	LJ	Liechtenstein	SI	Slovenia
CI	Côte d'Ivoire	LK	Sri Lanka	SK	Slovakia
CM	Cameroon	LR	Liberia	SN	Senegal
CN	China	LT	Lithuania	SZ	Swaziland
CS	Czechoslovakia	LU	Luxembourg	TD	Chad
CZ	Czech Republic	LV	Latvia	TG	Togo
DE	Germany	MC	Monaco	TJ	Tajikistan
DK	Denmark	MD	Republic of Moldova	TT	Trinidad and Tobago
EE	Estonia	MG	Madagascar	UA	Ukraine
ES	Spain	ML	Mali	UG	Uganda
FI	Finland	MN	Mongolia	US	United States of America
FR	France	MR	Mauritania	UZ	Uzbekistan
GA	Gabon			VN	Viet Nam

- 1 -

BRIDGED BIS-AMIDO GROUP 4 METAL COMPOUNDS
IN A CATALYST COMPOSITION

This invention relates to a novel catalyst composition and to its use in the oligomerisation and co-oligomerisation of one or more alkenes (olefins), in particular alpha olefins. More in particular, the invention relates to the oligomerisation of ethene to higher olefins, that is to olefins having 4-24 carbon atoms. The product higher olefins, in particular linear alpha olefins having 6-10 carbon atoms, are in great demand as intermediates in the preparation of detergents, lubricant additives and polyolefins. However, the oligomerisation reaction also produces less valuable products, such as internal olefins and branched olefins and olefins having a number of carbon atoms outside the range of 4-24. By further processing, these latter products can be converted to the desired linear alpha olefins.

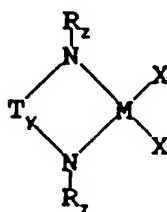
Polymerisation processes of olefins, such as the production of polyethylene from ethene, whereby homogeneous catalyst systems of the Ziegler-Natta type are used, are well known. Oligomerisation processes of lower olefins to higher olefins are also well known. In these processes use is made of varying catalyst systems. For example, from GB-A-135873 it is known that C₄-C₂₀ linear alpha olefins can be prepared from ethene by oligomerisation in the presence of a catalyst composition comprising a divalent nickel salt, a boron hydride and a tertiary organophosphorus compound. In WO 94/25416 a catalyst system for the preparation of C₄-C₂₄ linear alpha olefins is disclosed, comprising a bis-tetramethylcyclopentadienyl metallocene (the metal being

- 2 -

preferably zirconium) and a bulky and labile anion which is substantially non-coordinating (preferably one containing a plurality of boron atoms).

5 Bridged bis-amido Group 4 (IUPAC 1988 notation) metal compounds are also known, in catalyst systems for the preparation of polyolefins such as polyethylene and polypropylene.

10 In WO 92/12162 there are disclosed catalyst systems for the polymerisation of alpha olefins, comprising as a first component an amido transition metal compound of the general formula (I)



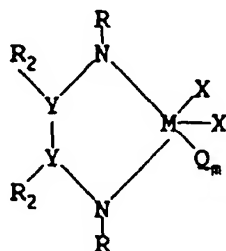
wherein M is zirconium, hafnium or titanium, N is a nitrogen atom having three substituents, X is any univalent anionic ligand, R is a hydrocarbyl, T is a covalent hydrocarbyl bridging group containing a Group IV-A or VI-A element (Deming notation; corresponding to Group 14 and 16 of the IUPAC 1988 notation) such as a silicon radical, y is 1 or 0 and z is 2-y, and as a second component alumoxane. The disclosed effect of this group of catalysts, which does not contain those having in the first component a disilyl, digermyl or distibyl bridging group and those having a second component different from alumoxane, is the production of solid stereoregular polyolefins having a molecular weight well in excess of 100,000.

25 There has now been found a novel catalyst composition based on bridged bis-amido Group 4 metal compounds which is particularly effective in the oligomerisation of lower

- 3 -

olefins to produce linear alpha-olefins. Compared to the best catalyst compositions based on metallocenes, this novel composition is distinguished by a highly selective production of the desirable linear alpha-olefins, with high turnover rates under very mild reaction conditions of temperature and pressure. The composition is further distinguished by the ease of preparation of the bis-amido Group 4 metal component from readily available pre-cursors.

The present invention provides a catalyst composition comprising a compound represented by the general formula (II)



wherein M is zirconium, hafnium or titanium; each N is a three coordinate nitrogen atom; each Y is Si, Ge or Sn; each X is, independently, a hydride or R; each R is, independently, a hydrocarbonyl one or more carbon atoms of which may be substituted by an element selected from Si, O, P, N and S; Q is a neutral Lewis base and m is a number from 0 to 2; or a dimer thereof; and a second component.

Preferably each R is independently chosen from the group of C₁-C₁₀ alkyl, C₆-C₁₅ aryl, C₃-C₁₀ cycloalkyl, alkylaryl or Si(R¹)₃; each R¹ being independently chosen from the group of C₁-C₁₀ alkyl, C₆-C₁₅ aryl, C₃-C₁₀ cycloalkyl or alkylaryl.

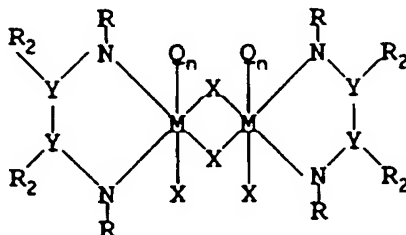
Preferably, Y is silicon.

Examples of the neutral Lewis base Q are diethylether, tetrahydrofuran, dimethylanilin, anilin, n-

- 4 -

butylamine and trimethylphosphine.

The dimer of the compound of formula (II) is according to the general formula (III)



wherein the symbols are as defined above, n being 0 or 1.

- 5 Preferred compounds of formula (II) according to the invention are {1,2-bis(t-butylamide)-tetramethyl-disilane}metal dibenzyl, {(Me₂SiNCMe₃)₂}M(CH₂Ph)₂ {1,2-bis(t-butylamide)-tetramethyldisilane}metal dimethyl, {(Me₂SiNCMe₃)₂}MMe₂
- 10 {1,2-bis(t-butylamide)-tetramethyldisilane}metal di(n-butyl), {(Me₂SiNCMe₃)₂}M(n-Bu)₂ {1,2-bis(t-butylamide)-tetramethyldisilane}metal diphenyl, {(Me₂SiNCMe₃)₂}MPh₂ {1,2-bis(t-butylamide)-tetramethyldisilane}metal di(4-
- 15 methylphenyl), {(Me₂SiNCMe₃)₂}M{CH₂(4-Me-Ph)}₂ {1,2-bis(t-butylamide)-tetramethyldisilane}metalla-cyclobutane, {(Me₂SiNCMe₃)₂} {MCH₂CH₂CH₂}
- 20 {1,2-bis(t-butylamide)-tetramethyldisilane}metal dihydride, {(Me₂SiNCMe₃)₂}MH₂ {1,2-bis(t-amylamide)-tetrametyldisilane}metal dibenzyl, {(Me₂SiNCMe₂Et)₂}M(CH₂Ph)₂ {1,2-bis(cyclohexylamide)-tetramethyldisilane}metal
- 25 dibenzyl, {(Me₂SiNCy)₂}M(CH₂Ph)₂ {1,2-bis(ethylamide)-tetramethyldisilane}metal dibenzyl, {(Me₂SiNEt)₂}M(CH₂Ph)₂

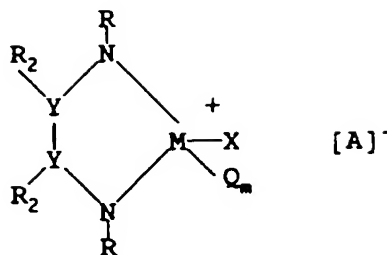
- 5 -

- {1,2-bis(phenylamide)-tetramethyldisilane}metal dibenzyl,
 { (Me₂SiNPh)₂ }M(CH₂Ph)₂
 {1,2-bis(2,6-dimethylphenylamide)-tetramethyldisilane}-
 metal dibenzyl,
 5 { (Me₂SiN[2,6-Me₂-Ph])₂ }M(CH₂Ph)₂
 {1,2-bis(trimethylsilylamide)-tetramethyldisilane}metal
 dibenzyl, { (Me₂SiNSiMe₃)₂ }M(CH₂Ph)₂
 {1,2-bis(tri(t-butyl)silylamide)-tetramethyldisilane}-
 metal dibenzyl, [{ (Me₂SiNSi(CMe₃)₃)₂ }M(CH₂Ph)₂
 10 {1,2-bis(t-butylamide)-tetraethyldisilane}metal dibenzyl,
 { (Et₂SiNCMe₃)₂ }M(CH₂Ph)₂
 {1,2-bis(t-butylamide)-tetraethyldisilane}metal dimethyl,
 { (Et₂SiNCMe₃)₂ }MMe₂
 {1,2-bis(t-butylamide)-tetraphenyldisilane}metal
 15 dibenzyl, { (Ph₂SiNCMe₃)₂ }M(CH₂Ph)₂
 {1,2-bis(t-butylamide)-tetramethyldigermane}metal
 dibenzyl, { (Me₂GeNCMe₃)₂ }M(CH₂Ph)₂, and
 {1,2-bis(t-butylamide)-tetramethyldistannane}metal
 dibenzyl, { (Me₂SnNCMe₃)₂ }M(CH₂Ph)₂
 20 in which the metal (M) is zirconium, hafnium or titanium.

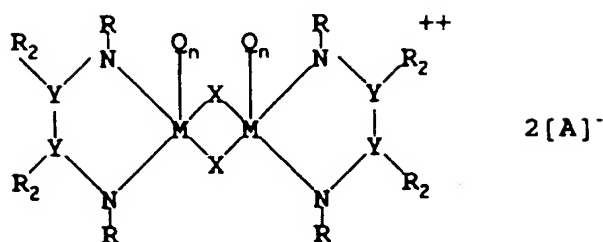
The present invention further comprises a catalyst
 composition comprising a first component which is a
 compound according to formula (II) or a dimer thereof,
 and a second component which is capable of providing a
 25 bulky and labile anion [A]⁻ which anion is substantially
 non-coordinating under the reaction conditions and
 contains at least one boron atom.

The first and second component together form an ionic
 compound of the general formula (IV)

- 6 -



or, of the general formula V



wherein the symbols are as defined above.

Examples of anion $[A]^-$ containing one boron atom are the borates of the general formula $[B(R^2)_4]^-$, wherein R^2 is a hydride, C_1 - C_{10} alkyl, C_6 - C_{15} aryl, C_3 - C_{10} cyclo-alkyl or alkylaryl, any of which can be substituted by one or more halogens, such as $[B(C_6F_5)_4]^-$, $[R^2B(C_6F_5)_3]^-$, $[B(FC_6H_4)_4]^-$, $[R^2B(FC_6H_4)_3]^-$, $[B\{(CF_3)_2(C_6H_3)\}_4]^-$ and $[R^2B\{(CF_3)_2(C_6H_3)\}_3]^-$.

Examples of anion $[A]^-$ containing a plurality of boron atoms are the carborates, such as $[B_{11}CH_{12}]^-$.

The second component can itself be an ionic compound of an anion $[A]^-$ as defined above and a cation. The cation is suitably a proton-donating cation, preferably a quaternary ammonium cation, in particular a trialkyl-ammonium cation such as tri-*n*-butylammonium, or dimethylanilinium. Alternatively a cation may be used in the second component which is not proton-donating, such as a metal cation e.g. a silver ion, or a triphenyl carbenium ion.

- 7 -

The second component can also be a neutral strongly Lewis acidic compound which is capable of abstracting one of the radicals X of the first component, thereby also contributing an anion $[A]^-$ as defined above.

- 5 Preferred second components in the catalytic composition according to the invention are the ionic compounds
- dimethylanilinium tetrakis(pentafluorophenyl)borate,
[PhMe₂NH][B(C₆F₅)₄]
- 10 tri(n-butyl)ammonium tetrakis(pentafluorophenyl)borate,
[Bu₃NH][B(C₆F₅)₄]
- dimethylanilinium tetrakis(2,3,5,6-tetrafluorophenyl)borate, [PhMe₂NH][B(2,3,5,6-C₆F₄H)₄]
- dimethylanilinium tetrakis(3,5-bis-trifluoromethylphenyl)borate, [PhMe₂NH][B(3,5-(CF₃)₂-C₆H₃)₄]
- 15 dimethylanilinium tetrakis(4-fluorophenyl)borate,
[PhMe₂NH][B(4-C₆H₄F)₄]
- dimethylanilinium tetraphenylborate, [PhMe₂NH][B(C₆H₅)₄]
- triphenylcarbonium tetrakis(pentafluorophenyl)borate,
20 [Ph₃C][B(C₆F₅)₄]
- ferrocenium tetrakis(pentafluorophenyl)borate,
[(C₅H₅)₂Fe][B(C₆F₅)₄]
- silver tetrakis(pentafluorophenyl)borate, [Ag][B(C₆F₅)₄]
- and
- 25 tri(n-butyl)ammonium 1-carbodecaborate, [Bu₃NH][CB₁₁H₁₂];
- and the neutral, strongly Lewis acidic compounds
- tris(pentafluorophenyl)borane, B(C₆F₅)₃
- tris(2,3,5,6-tetrafluorophenyl)borane, B(2,3,5,6-C₆F₄H)₃,
- and
- 30 trimethylboron, BMe₃.

The compounds of formula (II) according to the invention can be prepared in four steps.

- In a first step a compound Cl-YR₂YR₂-Cl (R and Y as defined above) is contacted with at least a four-fold
- 35 excess of a suitable amine RNH₂, followed by separation

- 8 -

of the RNH_3Cl by-product and removal of the solvent to give a bis-amine product $\text{RNHYR}_2\text{YR}_2\text{NHR}$.

5 In a second step the bis-amine product is contacted with a Group 1 metal derivative or with a Grignard derivative of a hydrocarbyl reagent, followed by removal of the solvent to give the Group 1 metal derivative or Grignard derivative of a bis-amide compound $\text{RNYR}_2\text{YR}_2\text{NR}$.

10 In a third step a Group 4 metal reactant is contacted in a suitable solvent with the Group 1 metal derivative or Grignard derivative of the bis-amide compound, followed by separation of the salt by-product and purification of the corresponding bis-amide Group 4 metal product.

15 And in a fourth step the bis-amide Group 4 metal product is contacted in a suitable solvent with a Group 1 metal derivative or with a Grignard derivative of a hydrocarbyl reagent, followed by salt removal and purification of the final compound of formula (II).

20 Exemplary reactions to prepare boron containing compounds which are suitable for use as the second component are the following.

- The neutral, strongly Lewis acidic $\text{B}(\text{C}_6\text{F}_5)_3$ can be obtained from the reaction of BCl_3 with LiC_6F_5 , as described by A.G. Massey and A.J. Park, J. Organomet. Chem. 2 1 64 245.

25 - The ionic $[\text{Me}_2\text{PhNH}]^+[\text{B}(\text{C}_6\text{F}_5)_4]^-$ can be prepared from $\text{B}(\text{C}_6\text{F}_5)_3$ and LiC_6F_5 , as also described in the Massey and Park publication, to give $\text{Li}(\text{OEt}_2)_n^+[\text{B}(\text{C}_6\text{F}_5)_4]^-$ - which can be converted by reaction with Me_2PhNHCl in dichloromethane to give $[\text{Me}_2\text{PhNH}]^+[\text{B}(\text{C}_6\text{F}_5)_4]^-$.

30 - $[\text{Ph}_3\text{C}]^+[\text{B}(\text{C}_6\text{F}_5)_4]^-$ can be prepared by reaction of $\text{Li}(\text{OEt}_2)_n^+[\text{B}(\text{C}_6\text{F}_5)_4]^-$ with Ph_3CCl in dichloromethane.

35 The catalyst composition may be formed by mixing together the two components, preferably in a solution in a suitable non-polar solvent such as toluene, benzene,

- 9 -

chlorobenzene, an alkane or an alkene, to form a liquid catalyst system. The two components are generally employed in substantially equimolar amounts, although the molar ratio of the first component to the second component may vary within the range of from 0.1 to 5.0. Such a quantity of the catalyst system is usually employed in the reaction mixture as to contain from 10^{-1} to 10^{-7} gram atoms, in particular from 10^{-3} to 10^{-5} gram atoms, of the metal per mole of the olefin to be reacted.

The two-component catalyst composition may be formed prior to its introduction to the reaction vessel, or it may be formed in situ.

Although not required for catalytic activity, further components may be added to the catalytic composition according to the invention, for example in order to increase the solubility and/or the stability of the composition. Organoaluminium compounds in relatively small amounts are efficient scavanging agents.

Examples of such organoaluminium compounds are trimethyl-aluminium, triethylaluminium, triisopropylaluminium, tri-isobutylaluminium, triphenyl-aluminium and diethylaluminium chloride.

The complete catalyst composition according to the invention can be used in solution. Alternatively, the catalyst composition can be loaded on a solid carrier, in particular an inorganic oxide such as silica, alumina, silica/alumina, titania, zirconia, magnesia and the like, but resinous support materials such as polyolefins can also be used. Suitable supports are the materials, composed of alumoxane and silica and marketed for example by WITCO GmbH, Bergkamen, Germany.

The invention also provides a process for oligomerising or co-oligomerising alpha olefins having a chain length of from 2 to 10 carbon atoms, especially involving ethene, to linear alpha olefins having a chain

- 10 -

length of from 4 to 24 carbon atoms, characterised in that the oligomerisation reaction is performed in the presence of the catalyst composition according to the invention.

5 The (co)oligomerisation reaction can be carried out in the liquid phase. When the catalyst compositions are loaded on an inert carrier the reaction is heterogeneous and can also be carried out in the gas phase. The reaction can be carried out in batch or in continuous
10 operation.

 The (co)oligomerisation reaction is generally, although not necessarily, carried out in an inert liquid which is suitably also the solvent for the catalyst components. The reaction is suitably carried out at a
15 moderate temperature, preferably in the range of -20 - +150 °C, more preferably at +10 - +100 °C. The reaction is suitably carried out under conditions of moderately elevated pressure, preferably in the range of
20 from 100 to 10000 kPa, more preferably from 200 to 2000 kPa. The optimum conditions of temperature and pressure for a particular reaction system in order to maximise the yield of the desired linear alpha olefins can be readily established by those skilled in the art, but it has been found that conditions of between
25 20 - 70 °C and between 200 - 1000 kPa are particularly advantageous in this respect with the catalyst systems of the present invention.

 The conditions of temperature and pressure are preferably selected to yield a product slate with a
30 'K factor' within the range of from 0.3 to 0.8. The K factor, which is indicative of the relative proportions of the product olefins, is the molar ratio of $[C_{n+2}]/[C_n]$ as calculated from the slope of the graph of $\log [C_n \text{ mol\%}]$ versus n , where n is the number of carbon
35 atoms in the particular product olefin.

- 11 -

The starting reactants can be supplied to the reactor together with an inert diluent, such as nitrogen or helium when the reactant is gaseous, and a liquid solvent, e.g. the same solvent as that of the catalyst components when the reactant is in the liquid form.

The reaction is preferably carried out in the absence of air or moisture.

Reaction times of from 1 minute to 5 hours have been found to be suitable, depending on the activity of the catalyst system and on the reaction conditions. After a suitable reaction time, a conventional catalyst deactivating agent such as water, methanol, or another alcohol may be added if desired to the reaction mixture in order to terminate the reaction. Alternatively, the reaction can simply be terminated by the introduction of air.

The product mixed olefins are preferentially linear alpha olefins having a chain length within the range of 4 to 24 carbon atoms, of which those having between 6 and 10 carbon atoms in the chain are particularly preferred. They may be suitably recovered by distillation and separation techniques known in the art.

If desired, unconverted starting material and oligomeric products having a molecular weight outside the desired molecular weight may be recovered, processed if necessary and recycled to be used as starting material in a subsequent oligomerisation reaction.

The invention will be further illustrated by the following examples.

EXAMPLES

1. Preparation of Complexes

1.1 General

All reactions were performed under nitrogen in a Braun MB 200-G dry box or under argon using standard Schlenk techniques. Solvents were dried by refluxing over

- 12 -

and distilling from standard reagents. 1,2-Me₄Si₂Cl₂ and H₂C(SiMe₂Cl)₂ were obtained from the reaction of Me₃SiCl/AlCl₃ with Si₂Me₆ and H₂C(SiMe₃)₂, respectively. PhMe₂NHCl was formed by reaction of PhNMe₂ and HCl.

5 Mg(CH₂Ph)Cl was obtained from the reaction of PhCH₂Br and Mg; Mg(CH₂Ph)₂(dioxane)_{0.5} was formed on reaction of Mg(CH₂Ph)Cl with dioxane. Si₂Me₆, H₂C(SiMe₃)₂, Me₂SiCl₂, t-BuNH₂, C₆F₅Br, BCl₃, PhNMe₂ and Ph₃CCl were obtained from Aldrich. Ether solutions of MeLi and and hexane
10 solutions of n-BuLi were obtained from Janssen (Belgium). All new products were characterized by NMR spectroscopy.

1.2 First Component of the Catalyst Composition

1.2.1 Preparation of {1,2-bis(t-butylamide)tetramethyl-disilane}-zirconium dibenzyl (Compound 1)

15 (a) 1,2-Bis(t-butylamine)tetramethyldisilane, (Me₂SiNHCM₃)₂

To a vigorously stirred solution of 1,2-Me₄Si₂Cl₂ (20.53 g, 110 mmol) in hexane (250 ml) was added a
20 solution of t-BuNH₂ (32.9 g, 450 mmol) in hexane (70 ml) using a dropping funnel. On completion of the addition, the white precipitate was removed by filtration, and washed with hexane (500 ml). The supernatant was reduced to dryness giving 27.2 g of (Me₂SiNHCM₃)₂ (95% yield).

25 (b) Dilithium 1,2-bis(t-butylamide)tetramethyldisilane, (Me₂SiNLiCMe₃)₂

To a stirred solution of (Me₂SiNHCM₃)₂ (14.2 g, 54.5 mmol) in hexane (20 ml) was added, via a dropping funnel, 68 ml of a 1.6 M n-BuLi solution in hexane (109 mmol). The crude (Me₂SiNLiCMe₃)₂ was obtained as a white solid
30 by removal of the solvent in vacuo.

(c) {1,2-Bis(t-butylamide)tetramethyldisilane}zirconium dichloride, {(Me₂SiNCMe₃)₂}ZrCl₂(THF)

To a solution of (Me₂SiNHCM₃)₂ (30.0 g, 115 mmol) in hexane (200 ml) was added 151 ml of a 1.6 M solution of
35 n-BuLi in hexane (242 mmol). The reaction was stirred for

- 13 -

45 min. and the solvent was then removed *in vacuo*. Toluene (150 ml) was added to the white solid and the resulting suspension introduced (via a transfer tube) into a Schlenk tube containing a vigorously stirred suspension of ZrCl_4 (26.8 g, 115 mmol) in toluene (150 ml). The reaction was allowed to warm slowly to room temperature and then stirred for 16 h. The solvent was removed *in vacuo* and the residue extracted with dichloromethane (2 x 250 ml). The extract was reduced to dryness and extracted with a dichloromethane/THF mixture (10:1). The extract was again reduced to dryness and extracted with toluene. Colourless microcrystalline $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{ZrCl}_2(\text{THF})$ (19.1 g, 40 mmol, 35% yield) was obtained on cooling the toluene solution to -20°C , separation of the crystals by filtration, washing with toluene and drying *in vacuo*.

(d) Compound 1, $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{Zr}(\text{CH}_2\text{Ph})_2$

Toluene (20 ml) at -78°C was added to a Schlenk tube at -78°C containing a mixture of $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{ZrCl}_2(\text{THF})$ (1.72 g, 3.45 mmol) and $\text{Mg}(\text{CH}_2\text{Ph})_2(\text{dioxane})_{0.5}$ (1.28 g, 5.18 mmol). The mixture was allowed to warm to room temperature and stirred for 90 min. The solvent was then removed *in vacuo*, the residue extracted with pentane, and the extract reduced to dryness. Cooling a pentane solution of the crude product to -40°C afforded 1.50 g (92 % yield) of pure yellow crystalline $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{Zr}(\text{CH}_2\text{Ph})_2$ (1).

1.2.2 Preparation of (1,2-bis(t-butylamide)tetramethyldisilane)-zirconium dimethyl, $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{ZrMe}_2$ (Compound 2)

To a suspension of $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{ZrCl}_2(\text{THF})$ (1.90 g, 4.5 mmol) in ether (50 ml) at -78°C in a Schlenk tube was added by syringe 5.8 ml of a 1.6 M solution of MeLi in ether (9.3 mmol). The mixture was allowed to warm slowly to room temperature and the solvent was then

- 14 -

removed *in vacuo*. The residue was then extracted with a small amount of pentane, and the solution cooled to -40°C , to give 0.87 g (51 % yield) of pure colourless crystalline $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{ZrMe}_2$ (2).

5 **1.2.3 Preparation of (1,2-bis(t-butylamide)tetramethyldisilane)-hafnium dibenzyl (Compound 3)**

(a) (1,2-Bis(t-butylamide)tetramethyldisilane)hafnium dichloride, $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{HfCl}_2(\text{THF})$

10 To a suspension of HfCl_4 (4.0 g, 12.5 mmol) in 40 ml of toluene/ether (10:1) was added via a dropping funnel, a solution of $(\text{Me}_2\text{SiNLiCMe}_3)_2$ (3.4 g, 12.5 mmol) in toluene/ether (20 ml; 10:1). The reaction was allowed to warm slowly to room temperature and then stirred for 16 h. The solvent was removed *in vacuo* and the residue
15 extracted with a dichloromethane/THF mixture (10:1). Reduction of the extract to dryness afforded 5.6 g (81% yield) of crude product. A sample of colourless crystalline $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{HfCl}_2(\text{THF})$ was obtained on cooling a toluene solution.

20 (b) Compound 3, $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{Hf}(\text{CH}_2\text{Ph})_2$

 To an ether solution of $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{HfCl}_2(\text{THF})$ (0.50 g, 1.14 mmol) at -78°C was added, by syringe, 1.37 ml of a 2.5 M solution of $\text{Mg}(\text{CH}_2\text{Ph})\text{Cl}$ (3.43 mmol). The mixture was allowed to warm to room temperature, stirred
25 for 1 h, and then reduced to dryness. Extraction with pentane afforded 0.43 g of crude product (71% yield). Crystallization from pentane at -40°C afforded pure pale yellow $\{(\text{Me}_2\text{SiNCMe}_3)_2\}\text{Hf}(\text{CH}_2\text{Ph})_2$ (3).

30 **1.2.4 Preparation of (bis(t-butylamide)dimethylsilane)zirconium dibenzyl (Compound 04)**

(a) Bis(t-butylamine)dimethylsilane, $\text{Me}_2\text{Si}(\text{NHCMe}_3)_2$

 Reaction of Me_2SiCl_2 (47.6 g, 370 mmol) with t-BuNH₂ (113.3 g, 1550 mmol) in hexane (700 ml) at room temperature, followed by brief reflux (10 min), cooling
35 of the mixture, filtration, and solvent removal from the

- 15 -

supernatant afforded an oily product. The $\text{Me}_2\text{Si}(\text{NHCMe}_3)_2$ was obtained in 98% purity by fractional vacuum distillation.

(b) Dilithium bis(t-butylamide)dimethylsilane,
5 $\text{Me}_2\text{Si}(\text{NLiCMe}_3)_2$

To a stirred solution of $\text{Me}_2\text{Si}(\text{NHCMe}_3)_2$ (14.7 g, 72 mmol) in hexane (200 ml) was added, via a dropping funnel, 93.4 ml of a 1.6 M n-BuLi solution in hexane (149 mmol). The mixture was then refluxed for 20 minutes,
10 cooled to room temperature, and the solvent removed *in vacuo*. Dissolution of the crude product in hexane and cooling to -40°C afforded 14.3 g (3 crops, 92% yield) of colourless crystalline $\text{Me}_2\text{Si}(\text{NLiCMe}_3)_2$.

(c) {Bis(t-butylamide)dimethylsilane}zirconium
15 dichloride, $\{\text{Me}_2\text{Si}(\text{NCMe}_3)_2\}\text{ZrCl}_2(\text{THF})_n$

Toluene (150 ml) at -78°C was added to a mixture of $\text{Me}_2\text{Si}(\text{NLiCMe}_3)_2$ (4.01 g, 18.7 mmol) and $\text{ZrCl}_4(\text{THF})_2$ (7.04 g, 18.7 mmol) at -78°C and the mixture allowed to warm to room temperature and further stirred for 90 min.
20 The solvent was removed *in vacuo* and the residue extracted with a mixture of THF and hexane.

The extract was reduced to dryness and then once again extracted with a THF/hexane mixture. Colourless microcrystalline $\{\text{Me}_2\text{Si}(\text{NCMe}_3)_2\}\text{ZrCl}_2(\text{THF})_n$ (3 crops, 4.5
25 g, ca. 50% yield) was obtained on cooling the solution to -40°C .

(d) Compound 04, $\{\text{Me}_2\text{Si}(\text{NCMe}_3)_2\}\text{Zr}(\text{CH}_2\text{Ph})_2$

Toluene (40 ml) at -40°C was added to a Schlenk tube at -40°C containing a mixture of $\{\text{Me}_2\text{Si}(\text{NCMe}_3)_2\}\text{ZrCl}_2(\text{THF})_n$ (0.83 g, 1.7 mmol) and $\text{Mg}(\text{CH}_2\text{Ph})_2(\text{dioxane})_{0.5}$ (0.50 g, 1.7 mmol). The mixture was allowed to warm to room temperature and then stirred for 30 min. 30 ml of a hexane/ether mixture (2:1) was added and the reaction
30 mixture was then filtered. The solvent was removed *in vacuo*, the residue extracted with hexane, and the extract
35

- 16 -

reduced to dryness.

Redissolution of the residue in hexane and cooling to -40 °C afforded 0.47 g (2 crops; 60% yield) of pure yellow crystalline $\{\text{Me}_2\text{Si}(\text{NCMe}_3)_2\}\text{Zr}(\text{CH}_2\text{Ph})_2$ (1).

5 **1.2.5 Preparation of [bis((t-butylamide)dimethylsilyl)methane]-zirconium dibenzyl (Compound 05)**

(a) Bis((t-butylamine)dimethylsilyl)methane, $\text{H}_2\text{C}(\text{Me}_2\text{SiNHCMe}_3)_2$

10 To a vigorously stirred solution of $\text{H}_2\text{C}(\text{SiMe}_2\text{Cl})_2$ (12.9 g, 64 mmol) in hexane (150 ml) was added a solution of t-BuNH₂ (19.4 g, 265 mmol) in hexane (90 ml) using a dropping funnel. On completion of the addition, the precipitate was removed by filtration and washed with hexane (100 ml). The solvent was removed from the
15 supernatant *in vacuo* giving 16.6 g of $\text{H}_2\text{C}(\text{Me}_2\text{SiNHCMe}_3)_2$ (94% yield).

(b) Dilithium bis((t-butylamide)dimethylsilyl)methane, $\text{H}_2\text{C}(\text{Me}_2\text{SiNLiCMe}_3)_2$

20 To a stirred solution of $\text{H}_2\text{C}(\text{Me}_2\text{SiNHCMe}_3)_2$ (8.30 g, 30 mmol) in hexane (25 ml) was added, via a dropping funnel, 38 ml of a 1.6 M n-BuLi solution in hexane (61 mmol). The crude solid $\text{H}_2\text{C}-(\text{Me}_2\text{SiNLiCMe}_3)_2$ was obtained by removal of the solvent *in vacuo*.

25 (c) [Bis((t-butylamide)dimethylsilyl)methane]zirconium dichloride, $\{\text{H}_2\text{C}(\text{SiMe}_2\text{NCMe}_3)_2\}\text{ZrCl}_2(\text{THF})$

To a suspension of ZrCl_4 (4.88 g, 20.9 mmol) in 50 ml of toluene/ether (10:1) was added, via a dropping funnel, a solution of $\text{H}_2\text{C}(\text{Me}_2\text{SiNLiCMe}_3)_2$ (6.0 g, 12.5 mmol) in toluene/ether (20 ml; 10:1). The reaction was allowed to
30 warm slowly to room temperature, then stirred for 16 h, and the volume was then reduced to 45 ml. The mixture was centrifuged and the supernatant reduced to dryness. The residue was extracted with a dichloromethane/THF mixture (20:1) and the extract reduced to dryness to yield 5.95 g
35 (56% yield) of $\{\text{H}_2\text{C}(\text{SiMe}_2\text{NCMe}_3)_2\}\text{ZrCl}_2(\text{THF})$.

- 17 -

(d) Compound 05, $\{H_2C(SiMe_2NCMe_3)_2\}Zr(CH_2Ph)_2$

Pentane (20 ml) at $-78^\circ C$ was added to a Schlenk tube at $-78^\circ C$ containing a mixture of $\{H_2C(SiMe_2NCMe_3)_2\}-ZrCl_2(THF)$ (0.80 g, 1.84 mmol) and $Mg(CH_2Ph)_2(dioxane)$ 0.5 (0.65 g, 2.58 mmol). The mixture was allowed to warm to room temperature and then stirred for 90 min. The solvent was removed *in vacuo* and the residue extracted with pentane (2 ml). On cooling the pentane solution to $-40^\circ C$, 0.36 g (36% yield) of pure yellow crystalline $\{H_2C(SiMe_2NCMe_3)_2\}-Zr(CH_2Ph)_2$ (05) was obtained.

1.3. Second Component of the Catalyst Composition

1.3.1 Tris(pentafluorophenyl)borane, $B(C_6F_5)_3$ (compound A)

$B(C_6F_5)_3$ (A) was obtained from the reaction of BCl_3 with LiC_6F_5 in hexane, following a literature procedure (A.G. Massey and A.J. Park, *J. Organomet. Chem.* 2 (1964) 245).

1.3.2 Dimethylanilinium tetrakis(pentafluorophenyl)-borate, $[Me_2PhNH]^+[B(C_6F_5)_4]^-$ (compound B)

$[Me_2PhNH]^+[B(C_6F_5)_4]^-$ (B) was synthesized by the metathesis reaction of $Li(OEt_2)_n^+[B(C_6F_5)_4]^-$ with $Me_2PhNHCl$ in dichloromethane. Pure crystalline product was obtained by removal of the lithium chloride by filtration, crystallization of the crude residue from dichloromethane/hexane solution, and ether removal under vacuum ($65^\circ C$, 3h). {The $Li(OEt_2)_n^+[B(C_6F_5)_4]^-$ was obtained from the reaction of LiC_6F_5 with $B(C_6F_5)_3$ in hexane/ether solution (A.G. Massey and A.J. Park, *J. Organomet. Chem.* 2 (1964) 245)}.

1.3.3 Triphenylcarbonium tetrakis(pentafluorophenyl)-borate, $[Ph_3C]^+[B(C_6F_5)_4]^-$ (compound C)

$[Ph_3C]^+[B(C_6F_5)_4]^-$ (C) was formed by reaction of $Li(OEt_2)_n^+[B(C_6F_5)_4]^-$ with Ph_3CCl in dichloromethane. Pure crystalline product was obtained by removal of the lithium chloride by filtration, crystallization of the crude residue from dichloromethane/hexane solution, and

- 18 -

ether removal under vacuum (65 °C, 3h).

2. Oligomerisation Reactions

2.1 General

5 The toluene solvent used was dried by refluxing over and distilling from sodium/benzophenone.

2.2 Catalyst Preparation

10 Catalyst preparation was carried out under nitrogen in a Braun MB 200-G dry box. A mixture of the first component (0.1 or 0.2 mmol) and the second component (0.1 or 0.2 mmol) were placed in a septum bottle and cooled to -40 °C. Toluene (10 ml) at -40 °C was added to the stirred mixture, which was then warmed to room temperature over 5 min, giving a yellow solution. The solution was then used immediately for the following
15 step.

2.3 Procedure for Oligomerisation

Oligomerisations were performed in a mechanically stirred, jacket-cooled, 1-litre steel autoclave which, prior to use, was placed under vacuum at 70 °C overnight.
20 Toluene (190 ml) at 25 °C, to which a known amount of *n*-hexylbenzene (about 0.5 g) had been added as internal standard, was introduced into the autoclave under vacuum, followed by the catalyst solution. The autoclave was then immediately pressurised with ethylene (ca. 700 kPa) and
25 the temperature and pressure monitored throughout the reaction, whilst maintaining a constant ethylene pressure. The exotherm, a qualitative measure of the relative activity of the catalysts, was measured as the difference between the starting temperature and the
30 highest temperature reached in the oligomerisation.

The oligomerisation was terminated after a period of 10-30 minutes by rapidly relieving the pressure in the autoclave and decanting the product mixture using a tap in the base of the autoclave into a collection bottle.
35 Exposure of the mixture to air resulted in rapid

- 19 -

deactivation of the catalyst.

2.4 Work-up/analysis

After removal of a sample for GC analysis, the crude product mixture was filtered to remove polyethylene and then placed under vacuum to yield the non-volatile oligomer mixture.

The GC analysis afforded the yields of C₄-C₃₀ and hexene isomers and was carried out as follows. The oligomer product distribution was determined using a HP 5890 gas chromatograph with a CPSIL5 capillary column (50 m x 0.25 mm, methylsilicone stationary phase) using helium as carrier gas and flame ionisation detection. Response factors for the even linear *alpha*-olefins and for *cis*- and *trans*-2-hexene, *cis*- and *trans*-3-hexene, 2-ethyl-1-butene and hexylbenzene (internal standard) were determined using a standard calibration mixture. The yields of the C₄-C₃₀ olefins obtained from the GC analysis were used to calculate the K factor in oligomerisation (via regression analysis).

The turnover rate (TOR) in Table 1, measured in moles of ethylene per hour converted to C₄-C₃₀ oligomers, per mole of catalyst, is calculated from the formula:

$$\frac{[(\text{weight in g of C}_4\text{-C}_{30} \text{ } \alpha\text{-olefins})]/(\text{moles catalyst})}{\text{x (reaction time in hours) x 28}}$$
 (28 being the molecular weight of ethylene).

The TOR allows the relative activities of the different catalysts to be compared. The relative amounts of the different hexene isomers found from the GC analysis is a measure of the selectivity of the catalysts towards linear *alpha*-olefin formation.

After washing the polyethylene with toluene and then acetone, the recovered polymer was dried under vacuum at 70 °C for three days and the yield then measured. The relative yields of oligomers and PE give a measure of the selectivity of the catalyst towards oligomerisation

- 20 -

compared to polymerisation.

The results are presented in the following Table 1.

Table 1: Ethylene oligomerization with cationic bis-amide group 4 complexes

Expt	Cat	T (°C)	t (min)	p C ₂ ⁼ (kPa)	Yld C ₄ ⁼ -C ₃₀ ⁼ (g)	TOR C ₄ ⁼ -C ₃₀ ⁼ (h ⁻¹)	Yld 1-C ₆ ⁼ (g)	Yld PE (g)	K	C ₆ ⁼ : 1H* (%)	C ₆ ⁼ : c2H* (%)	C ₆ ⁼ : t2H* (%)	C ₆ ⁼ : 2EB* (%)
1	1A	25	30	720	17.97	6,420	2.60	1.32	0.70	99.70	0	0	0.30
		(0.20)	[+22]										
2	1A	25	15	710	15.67	11,190	2.21	1.75	0.68	99.84	0	0	0.16
		(0.20)	[+9]										
3	1A	25	30	710	11.36	8,110	1.59	0.79	0.70	99.71	0	0	0.29
		(0.10)	[+5]										
4	1A	25	30	300	17.87	6,380	2.58	0.37	0.71	99.67	0	0	0.33
		(0.20)	[+13]										
5	1A	25	30	150	9.27	3,310	1.03	0.14	0.76	99.38	0	0	0.62
		(0.20)	[+5]										
6	1A	50	30	570	4.10	1,460	0.55	0.57	0.69	98.50	0.12	0.04	1.34
		(0.20)	[+12]										
7	2A	25	30	720	8.98	3,210	1.03	7.66	0.77	99.00	0	0.04	0.96
		(0.20)	[+22]										
8	3A	25	30	720	0.68	240	0.15	0.02	0.60	97.30	0	0	2.30
		(0.20)	[+3]										

Notes

* 1H = 1-hexene; c2H = *cis*-2-hexene; t2H = *trans*-2-hexene; 2EB = 2-ethyl-1-butene

Table 1 (Cont'd)

Expt	Cat	T (°C)	t (min)	p C ₂ ⁼ (kPa)	Yld C ₄ ⁼ -C ₃₀ ⁼ (g)	TOR C ₄ ⁼ -C ₃₀ ⁼ (h ⁻¹)	Yld 1-C ₆ ⁼ (g)	Yld PE (g)	K	C ₆ ⁼ : 1H* (g)	C ₆ ⁼ : c2H* (g)	C ₆ ⁼ : t2H* (g)	C ₆ ⁼ : 2EB* (g)
9	1B (0.20) [+54]	25	30	710	185.9	66,400	28.26	2.33	0.72	99.61	0.11	0.05	0.23
10	1B (0.10) [+39]	25	10	720	41.9	89,810	6.47	1.09	0.71	99.90	0.02	0	0.08
11	2B (0.10) [+18]	25	10	720	21.8	46,710	2.51	0.48	0.72	99.93	0	0	0.07
12	3B (0.10) [+2]	25	10	720	1.7	3,580	0.32	0.10	0.65	99.50	0	0	0.50
13	1C (0.10) [+62]	25	10	720	141.2	302,600	18.28	3.90	0.73	98.98	0.29	0.12	0.61
14	04A (0.20) [+2]	25	30	720	-	-	-	0.07	-	-	-	-	-
15	05A (0.20) [+1]	25	30	740	-	-	-	0.08	-	-	-	-	-

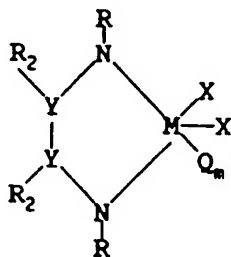
Notes

* 1H = 1-hexene; c2H = *cis*-2-hexene; t2H = *trans*-2-hexene; 2EB = 2-ethyl-1-butene

- 23 -

C L A I M S

1. A catalyst composition comprising a compound represented by the general formula (II)



wherein M is zirconium, hafnium or titanium; each N is a three coordinate nitrogen atom; each Y is Si, Ge or Sn; each X is, independently, a hydride or R; each R is, independently, a hydrocarbyl one or more carbon atoms of which may be substituted by an element selected from Si, O, P, N and S; Q is a neutral Lewis base and m is a number from 0 to 2; or a dimer thereof; and a second component.

2. A catalyst composition according to claim 1, wherein each R in the compound according to formula (II) is independently chosen from the group of C₁-C₁₀ alkyl, C₆-C₁₅ aryl, C₃-C₁₀ cycloalkyl, alkylaryl or Si(R¹)₃; each R¹ being independently chosen from the group of C₁-C₁₀ alkyl, C₆-C₁₅ aryl, C₃-C₁₀ cycloalkyl or alkylaryl.

3. A catalyst composition according to claim 1 or 2, characterized in that the compound according to formula (II) is chosen from the group of

(1,2-bis(t-butylamide)-tetramethyldisilane)metal

dibenzyl, {(Me₂SiNCMe₃)₂}M(CH₂Ph)₂

(1,2-bis(t-butylamide)-tetramethyldisilane)metal

dimethyl, {(Me₂SiNCMe₃)₂}MMe₂

(1,2-bis(t-butylamide)-tetramethyldisilane)metal di(n-

- 24 -

- butyl), $\{(Me_2SiNCMe_3)_2\}M(n-Bu)_2$
 $\{1,2-bis(t-butylamide)-tetramethyldisilane\}metal$
diphenyl, $\{(Me_2SiNCMe_3)_2\}MPh_2$
 $\{1,2-bis(t-butylamide)-tetramethyldisilane\}metal$ di(4-
5 methylphenyl),
 $\{(Me_2SiNCMe_3)_2\}M(CH_2(4-Me-Ph))_2$
 $\{1,2-bis(t-butylamide)-tetramethyldisilane\}metallacyclo-$
butane, $\{(Me_2SiNCMe_3)_2\}\{MCH_2CH_2CH_2\}$
 $\{1,2-bis(t-butylamide)-tetramethyldisilane\}metal$
10 dihydride, $\{(Me_2SiNCMe_3)_2\}MH_2$
 $\{1,2-bis(t-amylamide)-tetramethyldisilane\}metal$ dibenzyl,
 $\{(Me_2SiNCMe_2Et)_2\}M(CH_2Ph)_2$
 $\{1,2-bis(cyclohexylamide)-tetramethyldisilane\}metal$
dibenzyl, $\{(Me_2SiNCy)_2\}M(CH_2Ph)_2$
15 $\{1,2-bis(ethylamide)-tetramethyldisilane\}metal$ dibenzyl,
 $\{(Me_2SiNEt)_2\}M(CH_2Ph)_2$
 $\{1,2-bis(phenylamide)-tetramethyldisilane\}metal$ dibenzyl,
 $\{(Me_2SiNPh)_2\}M(CH_2Ph)_2$
 $\{1,2-bis(2,6-dimethylphenylamide)-tetramethyl-$
20 $disilane\}metal$ dibenzyl, $\{Me_2SiN[2,6-Me_2-Ph]\}_2M(CH_2Ph)_2$
 $\{1,2-bis(trimethylsilylamide)-tetramethyldisilane\}metal$
dibenzyl, $\{(Me_2SiNSiMe_3)_2\}M(CH_2Ph)_2$
 $\{1,2-bis\{tri(t-butyl)silylamide\}-tetramethyldisilane\}-$
metal dibenzyl, $\{[(Me_2SiNSi(CMe_3)_3)_2]M(CH_2Ph)_2$
25 $\{1,2-bis(t-butylamide)-tetraethyldisilane\}metal$ dibenzyl,
 $\{(Et_2SiNCMe_3)_2\}M(CH_2Ph)_2$
 $\{1,2-bis(t-butylamide)-tetraethyldisilane\}metal$ dimethyl,
 $\{(Et_2SiNCMe_3)_2\}MMe_2$
 $\{1,2-bis(t-butylamide)-tetraphenyldisilane\}metal$
30 dibenzyl, $\{(Ph_2SiNCMe_3)_2\}M(CH_2Ph)_2$
 $\{1,2-bis(t-butylamide)-tetramethyldigermane\}metal$
dibenzyl, $\{(Me_2GeNCMe_3)_2\}M(CH_2Ph)_2$, and
 $\{1,2-bis(t-butylamide)-tetramethyldistannane\}metal$
dibenzyl, $\{(Me_2SnNCMe_3)_2\}M(CH_2Ph)_2$
35 in which the metal (M) is zirconium, hafnium or titanium.

- 25 -

4. A catalyst composition according to claim 1 or 2, characterized in that Y is silicon.
5. A catalyst composition according to any one of claims 1-4, comprising a first component which is a compound according to formula (II) or a dimer thereof, and a second component which is capable of providing a bulky and labile anion $[A]^-$, which anion is substantially non-coordinating under the reaction conditions and contains at least one boron atom.
6. A catalyst composition according to claim 5, characterized in that the anion $[A]^-$ is a borate of the general formula $[B(R^2)_4]^-$, wherein R^2 is a hydride, C_1 - C_{10} alkyl, C_6 - C_{15} aryl, C_3 - C_{10} cycloalkyl or alkyl-aryl, any of which can be substituted by one or more halogens.
7. A catalyst composition according to claim 6, characterized in that the anion $[A]^-$ is chosen from the group of $[B(C_6F_5)_4]^-$, $[R^2B(C_6F_5)_3]^-$, $[B(FC_6H_4)_4]^-$, $[R^2B(FC_6H_4)_3]^-$, $[B\{(CF_3)_2(C_6H_3)\}_4]^-$ and $[R^2B\{(CF_3)_2(C_6H_3)\}_3]^-$.
8. A catalyst composition according to claim 5, characterized in that the anion $[A]^-$ is a carborate.
9. A catalyst composition according to claim 5, characterized in that the second component is chosen from the group of
- dimethylanilinium tetrakis(pentafluorophenyl)borate, $[PhMe_2NH][B(C_6F_5)_4]$
- tri(n-butyl)ammonium tetrakis(pentafluorophenyl)borate, $[Bu_3NH][B(C_6F_5)_4]$
- dimethylanilinium tetrakis(2,3,5,6-tetrafluorophenyl)-borate, $[PhMe_2NH][B(2,3,5,6-C_6F_4H)_4]$
- dimethylanilinium tetrakis(3,5-bis-trifluoromethyl-phenyl)borate, $[PhMe_2NH][B(3,5-(CF_3)_2-C_6H_3)_4]$
- dimethylanilinium tetrakis(4-fluorophenyl)borate, $[PhMe_2NH][B(4-C_6H_4F)_4]$

- 26 -

- dimethylanilinium tetrphenylborate, $[\text{PhMe}_2\text{NH}][\text{B}(\text{C}_6\text{H}_5)_4]$
triphenylcarbonium tetrakis(pentafluorophenyl)borate,
 $[\text{Ph}_3\text{C}][\text{B}(\text{C}_6\text{F}_5)_4]$
ferrocenium tetrakis(pentafluorophenyl)borate,
5 $[(\text{C}_5\text{H}_5)_2\text{Fe}][\text{B}(\text{C}_6\text{F}_5)_4]$
silver tetrakis(pentafluorophenyl)borate, $[\text{Ag}][\text{B}(\text{C}_6\text{F}_5)_4]$
tri(n-butyl)ammonium 1-carbodecaborate, $[\text{Bu}_3\text{NH}][\text{CB}_{11}\text{H}_{12}]$
tris(pentafluorophenyl)borane, $\text{B}(\text{C}_6\text{F}_5)_3$
tris(2,3,5,6-tetrafluorophenyl)borane, $\text{B}(2,3,5,6\text{-C}_6\text{F}_4\text{H})_3$,
10 and
trimethylboron, BMe_3 .
10. A catalyst composition according to any one of
claims 1-9, further comprising an organoaluminium
compound.
- 15 11. A catalyst composition according to any one of claims
1-10, which is loaded on a solid carrier chosen from the
group of silica, alumina, silica/alumina, titania,
zirconia, magnesia, or a resinous support material.
- 20 12. A process for oligomerising or co-oligomerising alpha
olefins having a chain length of from 2 to 10 carbon
atoms to linear alpha olefins having a chain length of
from 4 to 24 carbon atoms, characterized in that the
oligomerisation reaction is performed in the presence of
the catalyst composition according to any one of
25 claims 1-11.

INTERNATIONAL SEARCH REPORT

International Application No
PC/EP 96/01011

A. CLASSIFICATION OF SUBJECT MATTER IPC 6 B01J31/18 B01J31/12 B01J31/16 C07F7/10 C07F7/28		
According to International Patent Classification (IPC) or to both national classification and IPC		
B. FIELDS SEARCHED Minimum documentation searched (classification system followed by classification symbols) IPC 6 B01J C07F		
Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched		
Electronic data base consulted during the international search (name of data base and, where practical, search terms used)		
C. DOCUMENTS CONSIDERED TO BE RELEVANT		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	EP,A,0 332 970 (HOECHST AG) 20 September 1989	
A	WO,A,94 03271 (SMITHKLINE BEECHAM PLC; BELL DAVID (GB); MILLER DAVID (GB); ATTRIL) 17 February 1994	
A	US,A,3 487 096 (EISENHUTH WOLFGANG H ET AL) 30 December 1969	
A	FR,A,1 532 727 (STAUFER CHEMICAL COMPANY) 4 June 1968	
A	DE,A,20 23 433 (BRITISH TITAN PRODUCTS) 26 November 1970	
<input type="checkbox"/> Further documents are listed in the continuation of box C. <input checked="" type="checkbox"/> Patent family members are listed in annex.		
* Special categories of cited documents : "A" document defining the general state of the art which is not considered to be of particular relevance "E" earlier document but published on or after the international filing date "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "O" document referring to an oral disclosure, use, exhibition or other means "P" document published prior to the international filing date but later than the priority date claimed "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art "&" document member of the same patent family		
Date of the actual completion of the international search 3 July 1996		Date of mailing of the international search report 19.07.96
Name and mailing address of the ISA European Patent Office, P.B. 5818 Patentlaan 2 NL - 2280 HV Rijswijk Tel. (+ 31-70) 340-2040, Tx. 31 651 epo nl, Fax (+ 31-70) 340-3016		Authorized officer Thion, M

INTERNATIONAL SEARCH REPORT

International Application No

PC1/EP 96/01011

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP-A-0332970	20-09-89	DE-A- 3808868 CA-A- 1330984 JP-A- 1281149 US-A- 4965234	28-09-89 26-07-94 13-11-89 23-10-90
WO-A-9403271	17-02-94	AU-B- 4722393 CA-A- 2141790 CN-A- 1098415 EP-A- 0655950 JP-T- 8500099 ZA-A- 9305683	03-03-94 17-02-94 08-02-95 07-06-95 09-01-96 20-01-95
US-A-3487096	30-12-69	NONE	
FR-A-1532727	22-11-68	DE-A- 1643534 GB-A- 1189848 US-A- 3480653	01-07-71 29-04-70 25-11-69
DE-A-2023433	26-11-70	GB-A- 1246346	15-09-71